

Sheng-Qing Xia, Paul King and
Svilen Bobev*Department of Chemistry and Biochemistry,
University of Delaware, Newark, DE 19716,
USACorrespondence e-mail:
sbobev@chem.udel.edu

Key indicators

Single-crystal X-ray study
 $T = 120\text{ K}$
Mean $\sigma(\text{Mg-Zn}) = 0.001\text{ \AA}$
Disorder in main residue
 R factor = 0.011
 wR factor = 0.029
Data-to-parameter ratio = 11.8For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.Mg and Zn disorder in $\text{Mg}_{1.59(1)}\text{Zn}_{1.41(1)}\text{Sb}_2$

Single crystals of magnesium zinc diantimony, $\text{Mg}_{1.59(1)}\text{Zn}_{1.41(1)}\text{Sb}_2$, were grown by a reaction of elemental Mg, Zn and Sb in an excess of Zn acting as a flux. The compound crystallizes with the trigonal *anti*- La_2O_3 type (Pearson code hP5). The structure can be regarded as a substitutional derivative of α - Mg_3Sb_2 by replacing some of the Mg atoms in the tetrahedral position (site symmetry $3m$.) with Zn atoms. The other two atoms in the asymmetric unit also occupy special positions, *viz.* Mg ($\bar{3}m$.) and Sb ($3m$.)

Received 13 July 2006
Accepted 31 July 2006

Comment

A large number of intermetallic compounds with the *anti*- La_2O_3 type (also referred to as CaAl_2Si_2 type) are known (Villars & Calvert, 1991). The difference between the two is in the cations, which in the CaAl_2Si_2 structure are on Wyckoff position $1a$, whereas the La^{3+} ion in the La_2O_3 structure is located on one of the sites with symmetry $3m$. (Wyckoff position $2d$). The majority of the compounds crystallizing with

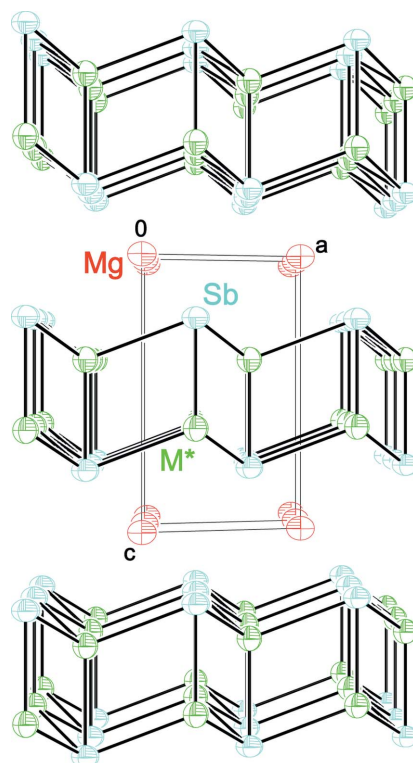


Figure 1
A projection of the crystal structure of $\text{Mg}_{1.59(1)}\text{Zn}_{1.41(1)}\text{Sb}_2$, viewed down the b axis. Displacement ellipsoids are drawn at the 90% probability level. Mg atoms are shown as red, the mixed Zn/Mg atoms (denoted as M^*) are shown as green, and the Sb atoms as blue ellipsoids. The unit cell is outlined.

this type are made up of main group elements only, although in recent years many isostructural transition metal phases have also been reported (Cordier & Schäfer, 1976; Zheng & Hoffmann, 1988; Bobev *et al.*, 2006). By means of the extended Hückel method, the formal electron count and the apparent requirement that the CaAl_2Si_2 type is favored over the rival ThCr_2Si_2 type for d^0 , d^5 and d^{10} configurations has been explained, at least semi-quantitatively (Zheng *et al.*, 1986; Burdett & Miller, 1990).

Fig. 1 shows a perspective view of the layered trigonal crystal structure of $\text{Mg}_{1.59}\text{Zn}_{1.41}\text{Sb}_2$, which can be readily described as double corrugated layers of $(\text{Zn,Mg})_2\text{Sb}_2^{2-}$, separated by Mg^{2+} cations. The corresponding interatomic distances are given in Table 1. The double-layers could be derived by puckering of 'dimerized' graphite-like layers or by splitting of a wurtzite-type framework and a subsequent reconstruction (Burdett & Miller, 1990). Further details on the local and extended bonding in CaAl_2Si_2 -type intermetallics have been given full consideration in earlier publications (Deller & Eisenmann, 1977; Zheng *et al.*, 1986; Bobev *et al.*, 2006).

The structure of the title compound can be viewed as a derivative of $\alpha\text{-Mg}_3\text{Sb}_2$ (Loehberg, 1934; Martinez-Ripoll *et al.*, 1974), in which almost 70% of the Mg atoms occupying the tetrahedral position with Wyckoff position $2d$ ($1/3, 2/3, z$) are replaced by Zn atoms. This is supported not only by refinement of the structure with unconstrained occupancy on that site but also by the corresponding distances, which are much shorter than those observed in the related AMg_2Sb_2 compounds ($A = \text{Mg, Ca, Sr, Ba, Eu}$ and Yb) (Deller & Eisenmann, 1977; Wartenberg *et al.*, 2002). Analogously, the refined distances are longer than those in AZn_2Sb_2 (Mewis, 1978), in good agreement with the Zn and Mg covalent radii (1.364 Å for Mg and 1.213 Å for Zn; Pauling, 1960). These experimental findings are well supported by the local bonding and site preference analyses (Zheng *et al.*, 1986; Burdett & Miller, 1990) based on band structure calculations.

Lastly, we point out that similar disorder of Mg and Zn has been proposed for the isostructural solid solution $\text{Mg}_{1.86}\text{Zn}_{1.14}\text{Sb}_2$ (Loehberg, 1934); however, this report erroneously suggests that Mg on both the octahedral and tetrahedral positions is replaced by Zn. Our refinement of $\text{Mg}_{1.59}\text{Zn}_{1.41}\text{Sb}_2$, which is the Zn-rich phase, unambiguously shows that the octahedral Mg position [Wyckoff symmetry $1a$, (0,0,0)] is not disordered.

Experimental

Handling of the pure metals and of the reaction mixtures was carried out either in an argon-filled glove-box or under vacuum. All starting materials were used as received: Mg (Alfa, turnings, 99.98% metal basis), Zn (Alfa, shot, 99.999%), Sb (Alfa, ingot, 99.99%). The synthesis was performed by loading the elements in a ratio Mg:Zn:Sb = 1:10:2 into an alumina crucible, which was subsequently sealed in an evacuated fused silica jacket. The reaction mixture was heated to 1273 K at a rate of 300 K h⁻¹, allowed to equilibrate at this temperature for 5 h, and cooled slowly to 873 K at a rate of 5 K h⁻¹. After reaching this temperature, it was removed from the furnace and

the excess of molten Zn was removed by centrifugation. Black, air-stable crystals of the title compound were the only product of this reaction.

Crystal data

$\text{Mg}_{1.59}\text{Zn}_{1.41}\text{Sb}_2$	$D_x = 5.180 \text{ Mg m}^{-3}$
$M_r = 374.32$	Mo $K\alpha$ radiation
Trigonal, $P\bar{3}m1$	$\mu = 18.11 \text{ mm}^{-1}$
$a = 4.3966$ (9) Å	$T = 120$ (2) K
$c = 7.171$ (3) Å	Irregular fragment, black
$V = 120.04$ (6) Å ³	$0.08 \times 0.06 \times 0.04 \text{ mm}$
$Z = 1$	

Data collection

Bruker SMART APEX diffractometer	1294 measured reflections
ω scans	118 independent reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 2003)	117 reflections with $I > 2\sigma(I)$
$T_{\min} = 0.287$, $T_{\max} = 0.482$	$R_{\text{int}} = 0.023$
	$\theta_{\max} = 26.1^\circ$

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_o^2) + (0.017P)^2]$
$R[F^2 > 2\sigma(F^2)] = 0.011$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.029$	$(\Delta/\sigma)_{\max} < 0.001$
$S = 1.29$	$\Delta\rho_{\max} = 0.37 \text{ e \AA}^{-3}$
118 reflections	$\Delta\rho_{\min} = -0.69 \text{ e \AA}^{-3}$
10 parameters	

Table 1

Selected bond lengths (Å).

Mg—Sb	3.0403 (6)	Zn1/Mg1—Sb	2.7740 (13)
Zn1/Mg1—Sb ⁱ	2.7470 (6)	Zn1/Mg1—Zn1/Mg1 ⁱⁱ	3.0685 (10)

Symmetry codes: (i) $-x + 1, -y + 2, -z + 1$; (ii) $-x, -y + 1, -z + 1$.

In the structure refinement, the full occupancies for all sites were verified by freeing the site occupation factor for an individual atom, while the remaining occupation parameters were kept fixed. This proved that the Sb site (site symmetry $3m.$) and one of the Mg sites (site symmetry $\bar{3}m.$) are fully occupied with deviations from full occupancy within 3σ . The second Mg site (site symmetry $3m.$) refined with occupancy much higher than 100%. When Zn was refined in that position, the occupancy factor was significantly below unity. Therefore, this position (labeled Zn1/Mg1) was modeled as a statistical mixture of Zn and Mg, and refined as 70.3 (4)% Zn and 29.7 (4)% Mg.

Data collection: SMART (Bruker, 2002); cell refinement: SAINT (Bruker, 2002); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2001); program(s) used to refine structure: SHELXTL; molecular graphics: XP in SHELXTL; software used to prepare material for publication: SHELXTL.

This work was funded by a University of Delaware start-up grant. The authors also thank the University of Delaware Research Foundation for financial support.

References

- Bobev, S., Merz, J., Lima, A., Fritsch, V., Thompson, J. D., Sarrao, J. L., Gillessen, M. & Dronskowski, R. (2006). *Inorg. Chem.* **45**, 4047–4054.
- Bruker (2002). SMART and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Burdett, J. K. & Miller, G. J. (1990). *Chem. Mater.* **2**, 12–26.
- Cordier, G. & Schäfer, H. (1976). *Z. Naturforsch. Teil B*, **31**, 1459–1461.
- Deller, K. & Eisenmann, B. (1977). *Z. Naturforsch. Teil B*, **32**, 612–616.

- Loehberg, K. (1934). *Z. Phys. Chem. B*, **27**, 381–403.
- Martinez-Ripoll, M., Haase, A. & Brauer, G. (1974). *Acta Cryst.* **B30**, 2006–2009.
- Mewis, A. (1978). *Z. Naturforsch. Teil B*, **33**, 382–384.
- Pauling, L. (1960). *The Nature of the Chemical Bond*. Ithaca, New York, USA: Cornell University Press.
- Sheldrick, G. M. (2001). *SHELXTL*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Sheldrick, G. M. (2003). *SADABS*. University of Göttingen, Germany.
- Villars, P. & Calvert, L. D. (1991). *Pearson's Handbook of Crystallographic Data for Intermetallic Compounds*, 2nd ed. Materials Park, Ohio, USA: American Society for Metals.
- Wartenberg, F., Kranenberg, C., Pocha, R., Johrendt, D., Mewis, A., Hoffmann, R.-D., Mosel, B. D. & Pöttgen, R. (2002). *Z. Naturforsch. Teil B*, **57**, 1270–1276.
- Zheng, C. & Hoffmann, R. (1988). *J. Solid State Chem.* **72**, 58–71.
- Zheng, C., Hoffmann, R., Nesper, R. & von Schnering, H. G. (1986). *J. Am. Chem. Soc.* **108**, 1876–1884.